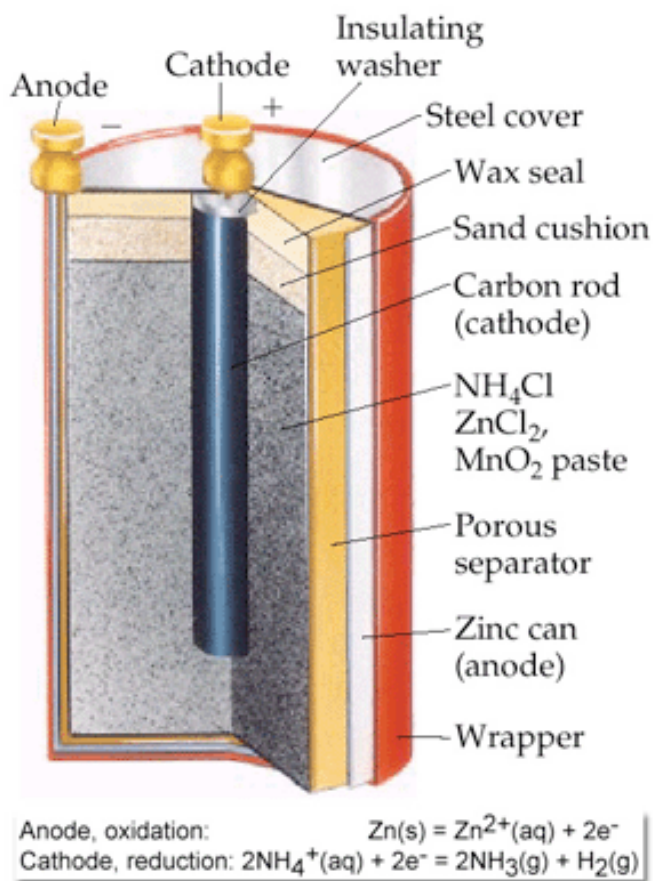


# Inside a Dry Cell

## How a dry cell works:

Many years ago, a man by the name of Alessandro Volta combined two different metals and some salty moist paper and layered them into a stack, and discovered that this unusual combination creates an electrical current.

Interestingly enough, pictured at the RIGHT, is technically called a DRY CELL. If you look the diagram, you will NOT find ACID inside a dry cell. However, a dry cell is not completely dry. It does have moist electrolyte inside it that is actually an Alkaline or base. This chemical is highly corrosive and begins a chemical reaction called "rusting".



The positive, or TOP of a Dry Cell, is the carbon rod. It goes down the center of the dry cell. It too has a chemical reaction that creates a positive charge.

A chemical reaction between the black paste inside and the zinc can makes the (-) electrons collect on the zinc can. When the positive and negative terminals of the dry cell are connected by means of the wire, electrons move through the wire from the negative terminal to the positive terminal. Electric current always moves from a negative terminal to a positive terminal.

Regardless of the size of the Dry Cell (D, C, AA, or even AAA) dry cells only produce 1.5 volts of electricity. The only difference is the actual length and width of the dry cell to fit different sized electronic devices. These types of dry cells are NOT RECHARGEABLE. Dry cells kept in the refrigerator tend to last longer. Rechargeable dry cells are made of metals like Lithium, Nickel, and Cadmium, which collect a charge in a different kind of chemical reaction.

A BATTERY is a combination of two or more cells that are stacked together. A car battery is called a WET CELL because it does contain liquid acid to create the chemical reaction.